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# Simple Chemical Reactions For The Preparation Of Metal Carbonates To Be Used As a Nucleus For The Preparation Of Nanometer Oxides: Infrared Spectral Studies

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**Abstract:** The work reported in this study deals with synthesis, characterization and structural investigation of some new compounds including novel methods for the synthesis of some metal carbonates. Metal(II) carbonate,  $MCO_3$ , is obtained by a new very simple synthetic method involving the reaction of aqueous solutions of some metal ions with urea at  $\sim 90^\circ C$ . The infrared spectrum of the formed product clearly indicates the absence of bands due to urea and shows the characteristic bands of carbonate ion. The  $(CO_3^{2-})$  ion is planar and therefore, it belongs to the  $D_{3h}$  symmetry. It is expected to display four modes of vibration,  $A_1 + A_2 + 2E$  ( $E$  is a doubly degenerate motion). The zinc(II) and cobalt(II) carbonates, respectively, were prepared using a new chemical procedure through the interaction of  $Zn(HCOO)_2 \cdot xH_2O$  and  $Co(HCOO)_2 \cdot xH_2O$  in aqueous solution with a urea as a simple organic precursor at  $\sim 90^\circ C$  for 12 hrs.

**Keywords:** carbonate  $CO_3^{2-}$ , infrared spectra, cobalt(II) nitrate, zinc(II) nitrate, urea.

## I. INTRODUCTION

Urea is physiologically very important. It is the chief nitrogenous product of protein metabolism. Urea has a melting point of  $132^\circ C$ , soluble in water and ethanol, but insoluble in ether. Urea is used for preparing formaldehyde-Urea resin (plastics) [1], barbiturates [2], and fertilizers [3-6]. Urea is also extensively used in the paper industry to soften cellulose and has been used to promote healing in infected wounds and many other applications in the field of medicine [7-9]. Some metal-urea complexes have pharmaceutical application, e.g., the platinum-urea complex which is used as antitumor [10]. Yamaguchi and Stewart [11, 12] were assigned all of the observed frequencies in the spectra of urea and urea- $d_4$ . The two vibrations of the frequencies at  $1686$  and  $1603\text{cm}^{-1}$  were assigned as the  $1686\text{cm}^{-1}$  band due to CO stretching vibration and the  $1603\text{cm}^{-1}$  band for  $NH_2$  bending motion. The calculations studied by Yamaguchi showed that for the band at  $1686\text{cm}^{-1}$ , the contribution of the  $NH_2$  bending motion is greater than that of CO stretching motion. The infrared bands of urea- $d_4$  observed at  $1245$  and  $1154\text{cm}^{-1}$  are assigned to  $ND_2$  bending vibrations. This assignment is consistent with the observed depolarization degrees of the Raman lines. The  $1464\text{cm}^{-1}$  frequency of urea is assigned to the CN stretching vibration. The corresponding frequency of urea- $d_4$  is observed at  $1490\text{cm}^{-1}$ . The  $1150\text{cm}^{-1}$  band is assigned to  $NH_2$  rocking vibrations. The reactions between transition metal ions and urea at room temperature have been studied extensively [13-17]. The infrared spectra of these complexes clearly indicated that urea molecule behaves as a mono dentate ligand and coordinates to the metal ions through the oxygen atom and not the nitrogen atom. The nature of the reaction products depend strongly on the type of metal ions and so the metal salt used. The novelty in our previously studies [18-27] were oriented to the reaction of urea ligand with different metals such as Co(II), pb(II), Sn(II), Cr(III), Fe(III), Au(III), Sn(IV), V(V) and Mo(IV) at high

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temperature which demonstrate that the types of metal ions beside their anions have a pronounced effect on the nature of the reaction products. The published papers were trended for the reaction of urea with different metal salts at elevated temperature lead to discovering a novel method for preparation  $\text{PbCO}_3$  and  $\text{CoCO}_3$  [21], lanthanide carbonates [23, 27], limonite,  $\text{FeO(OH)}$  [20],  $2\text{ZnCO}_3 \cdot 3\text{Zn(OH)}_2$  [19],  $\text{SnOCl}_2 \cdot 2\text{H}_2\text{O}$  [18], ( $\text{Cr}_2\text{O}_3$ ,  $\text{MnO}_2$ ,  $\text{MoO}_3$  and  $\text{WO}_3$ ) oxides resulted from a novel oxidation reduction reaction between ( $\text{K}_2\text{CrO}_4$  or  $\text{K}_2\text{Cr}_2\text{O}_7$ ),  $\text{KMnO}_4$ ,  $\text{Na}_2\text{MoO}_4$  and  $\text{Na}_2\text{WO}_4$ , respectively, with urea in an aqueous solution at  $\sim 85^\circ\text{C}$  [27]. The shining point in this paper was aimed to identify the reaction mechanisms of the products resulted during the reaction of urea with  $\text{Zn(HCOO)}_2 \cdot x\text{H}_2\text{O}$  and  $\text{Co(HCOO)}_2 \cdot x\text{H}_2\text{O}$  at  $90^\circ\text{C}$  for 12 hrs in aqueous media. The reaction products were isolated as solids and characterized by infrared spectroscopy.

## II. EXPERIMENTAL

### 2-1- Materials and synthesis

All chemicals used throughout this paper were analytical pure.  $\text{MCO}_3$  ( $\text{M} = \text{Co(II)}$  and  $\text{Zn(II)}$ ) were prepared by mixing an aqueous solutions (50 ml) of 0.1M of urea with 0.01M of the respective  $\text{Co(HCOO)}_2 \cdot x\text{H}_2\text{O}$  and  $\text{Zn(HCOO)}_2 \cdot x\text{H}_2\text{O}$ , respectively. The mixtures were heated at  $90^\circ\text{C}$  for 12 hrs in a hot plate. The solid products compounds were filtered off, washed several times with hot water, dried at  $90^\circ\text{C}$  in an oven. The yields of the obtained  $\text{Co(II)}$  and  $\text{Zn(II)}$  carbonates were varied in the range 55-to-60% depending upon the type of metal as well as on the counter ions associated with the metal ion.

### 2-1- Instruments

Carbonate content in the cobalt(II) and zinc(II) compounds were determined by dissolving a sample of each product in excess standard HCl and the excess of HCl was determined using standard sodium carbonate [28]. The percentage of cobalt(II) and zinc(II) in their compounds were determined gravimetrically method till constant weight and stable oxides formula. The infrared spectra of urea, all reactants and products were recorded using a Bruker FT-IR Spectrophotometer in Taif University spectroscopic Laboratory.

## III. RESULTS AND DISCUSSION

The reaction of aqueous solutions of urea with formate salts of cobalt(II) and zinc(II) at  $90^\circ\text{C}$  produces a pink and white solid powder products, respectively. The infrared spectra of urea as well as the reaction products of different cobalt(II) and zinc(II) salts with urea at high temperature less than boiling points of water solvent were obtained. The spectra of free urea ligand, cobalt(II) and zinc(II) carbonates are shown in Fig. 1 and 2, respectively. The band assignments for the products are given in Table 1. The infrared spectra show no bands due to any of the reactants and of coordinated urea, but instead, a group of bands characteristic for the ionic Carbonate,  $(\text{CO}_3)^{2-}$ , is appeared [29]. Based on this fact, the volumetric determination of  $(\text{CO}_3)^{2-}$  group with standard solution of HCl and beside that the infrared spectra of the commercially obtained  $\text{CoCO}_3$  and  $\text{ZnCO}_3$  are the same as that of the reaction products. The products obtained were identified as  $\text{CoCO}_3 \cdot x\text{H}_2\text{O}$  and  $\text{ZnCO}_3 \cdot x\text{H}_2\text{O}$ . The infrared assignments agree quite well with those known [29] for the ionic carbonate  $(\text{CO}_3)^{2-}$ . Previous studies [18-27] indicated that the nature of the reaction product obtained from the reaction of metal ions with urea at high temperature depends upon the type of metal ion and in some cases on the nature of the metal salts used. Oxalato urea complexes of manganese(II)  $[\text{MnU}_2(\text{C}_2\text{O}_4)]\text{H}_2\text{O}$  ( $\text{U}=\text{Urea}$ ) was prepared [30] from the reaction of  $\text{MnC}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$  and urea at pH  $\sim 3$  in 1:1, 1:2, 1:3 and 1:4 molar ratios. This complex was characterized by its IR-spectrum, X-ray diffraction, and thermal analysis. The data obtained indicates that, the oxalato group and the  $\text{H}_2\text{O}$  molecule are in the inner-sphere. The thermal analysis data indicate that, this complex  $[\text{MnU}_2(\text{C}_2\text{O}_4)]\text{H}_2\text{O}$  decomposes at  $400^\circ\text{C}$  in three stages to  $\text{Mn}_3\text{O}_4$  as the final product. Taking the  $[\text{MnU}_2(\text{C}_2\text{O}_4)]\text{H}_2\text{O}$  parent complex as a guide in our paper, so, at high temperature the role of  $\text{Co(II)}$  ions in decomposing the coordinated urea in the form of  $[\text{CoU}_2(\text{HCO}_2)_2]$  could be understood according to the following reactions;

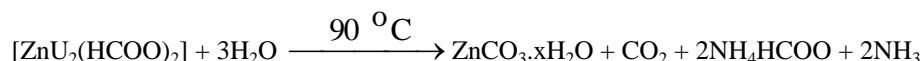
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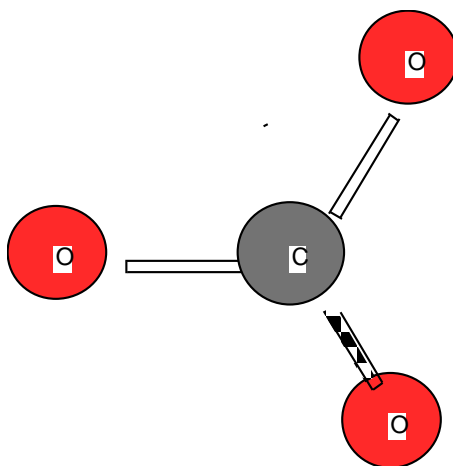
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Adducts of  $\text{ZnX}_2 \cdot 2\text{urea}$  ( $\text{X} = \text{Cl}, \text{Br}, \text{I}$ ),  $\text{CdX}_2 \cdot 2\text{urea}$  ( $\text{X} = \text{Br}, \text{I}$ ),  $\text{CdX}_2 \cdot \text{urea}$  ( $\text{X} = \text{Cl}, \text{Br}, \text{I}$ ) and  $\text{HgX}_2 \cdot \text{urea}$  ( $\text{X} = \text{Cl}, \text{Br}$ ) have been characterized [31]. The shifts of the stretching vibrations to low frequency ( $\nu\text{CO}$ ) and to high frequencies ( $\nu\text{CN}$  and  $\nu\text{NH}$ ) indicated that urea is oxygen-bonded to the metals. At high temperature the role of  $\text{Cd}(\text{II})$  ions in decomposing the coordination urea in the form of  $[\text{Zn}(\text{urea})_2(\text{HCOO})_2]$  could be understood as follows;



The  $(\text{CO}_3)^{2-}$  ion is planar and therefore, it belongs to the  $D_{3h}$  symmetry. It is expected to display four modes of vibrations,  $A'_1 + A''_2 + 2E'$  ( $E'$  is a doubly degenerate motion). The vibration  $A'_1$  is only Raman active while the other  $\nu_2$ ,  $\nu_3$  and  $\nu_4$  are infrared active. The stretching vibrations of the type;  $\nu(\text{C}-\text{O})$ ;  $\nu_3(E')$  is observed as a strong broad band in the range of  $1570\text{-}1330\text{ cm}^{-1}$  while the stretching vibration  $\nu(\text{C}-\text{O})$ ;  $\nu_1(A'_1)$  is observed in the region  $\sim 1070\text{ cm}^{-1}$  as a medium-to-weak band. It should be indicated here that this motion ( $A'_1$ ) should be infrared inactive, its weak appearance in the spectrum of  $\text{MCO}_3$  could be related to weak solid-solid interactions causes the symmetry of  $\text{CO}_3^{2-}$  to be lowered than  $D_{3h}$ . The out of plane of vibration  $\delta(\text{OCO})$ ;  $\nu_2(A''_2)$  is observed in the range of  $\sim 860\text{ cm}^{-1}$  as a strong band while, the angle deformation bending vibration  $\delta(\text{OCO})$ ;  $\nu_4(E')$  appear in the range of  $720\text{ cm}^{-1}$  as a medium strong. The infrared spectra of metal(II) carbonate,  $\text{MCO}_3$  (formula 1), show that, some these products are hygroscopic and its clearly have moisture water. The band related to the stretching vibration  $\nu(\text{O}-\text{H})$  of uncoordinated  $\text{H}_2\text{O}$  is observed as expected in the range of  $\sim 3000\text{ cm}^{-1}$  Table 1.



**Formula 1:** Structure of carbonate ions

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## REFERENCES

1. Feldman, D. and Barbalata, A. "Synthetic Polymers", Chapman & Hall, London (1996).
2. Finar, I. L. "Organic Chemistry", Longman group limited, London (1973) P. 460.
3. Rahman, M. J. , Bozadjiev , P. and Polovski, Y. fert. Res., 38(2), 89 (1994).
4. George, S., Chellapandian, M., Sivasankar, B. and Jayaraman, K. Bioprocess Eng., 16(2), 83 (1997).
5. Wang, X. J. and Douglas, L. A. Agrochimica, 40(5-6), 209 (1996).
6. Yerokun, O. A., S. Afr. J. plant soil, 14(2), 63 (1997).
7. Heinig, R. SOFW J., 122(14), 998 (1996).
8. Gnewuch, C. T. and Sosnovsky, G. Chem. Rev., 97(3), 829 (1997).
9. Miyagawa, C. I. Drug Intell. & Clin. Pharma., 20, 527 (1986).
10. Sugimura, M., Kameyama, Y., Hashimoto, T., Kobayashi, T. and Muramatsu, S. Chem. Abs., 112, 63 (1990).
11. Yamaguchi, A., Miyazawa, T., Shimanouchi , T. and Mizushima, S. Spectrochim. Acta, 10, 170 (1957).
12. Stewart, J. E. , J. Chem. Phys., 26, 248 (1957).
13. Srivastava, P.C. and Aravindaksaahn, C., Z. Phys. Chem. Leipzig, 264, 61 (1983)
14. Schafer , M.and Curran, C. , Inorg. Chem., 5, 256 (1966)
15. Zarli, B., Dall'Olio, G. and Sindellari,, L. J. Inorg. Nucl. Chem., 38, 491 (1976)
16. Barbier, J.P. , and Hugel, R.P. , J. Inorg. Nucl. Chem., 39, 2283 (1977)
17. Barbier, J.P. and Hugel, R. Inorg. Chimica Acta, 10, 93 (1974)
18. Teleb, S. M. , Refat, M. S. , Sadeek, S. A. , Ukrainian Chem. J., 68(5), 35 (2004).
19. Sadeek, S. A. , Refat, M. S. ,J. Coord. Chem., 58(18), 1727 (2005).
20. Sadeek, S. A. , Refat, M. S. , Teleb, S. M. , J. Korea, Chem. Soc., 48(4), 358 (2004).
21. Refat, M. S., Teleb, S. M. , Sadeek, S. A. , Spectrochim. Acta, A(60), 2803 (2004).
22. Refat, M. S. , Sadeek, S. A. , Teleb, S. M. , J. Argen. Chem. Soc., 92(4-6), 23 (2004).
23. Refat, M. S. , Synth. React. Inorg. Met.-Org. Chem., 34(9), 1605 (2004).
24. Nour, E. M. , Teleb, S.M. , Al-Khososy, N.A. , Refat, M. S. , Synth. React. Inorg. Met-Org. Chem., 27(4), 505 (1997).
25. Refat , M.S. and Sadeek, S.A. , Latvian J. Chem., 4, 343 (2005).
26. Refat, M.S. , Sadeek, S.A. and Nasr, D.E. Bull. Chem. Technolog. Maced., 24(2), 153 (2005).
27. Teleb, S.M. , and Refat, M.S. Bull. Chem. Technolog. Maced., 25(1), 57 (2006).
28. Vogel, "Qualitative Inorganic Analysis", John wiely & Sans, Inc. New York, (1987)
29. Nakamoto, K. "Infrared and Raman Spectra of inorganic and coordination compounds", 3rd edn (New York: Wiley Interscience) 1978.
30. Yu. Ya. Kharitonov, L. N. Ambroladze and L. M. Tkavadze, Koord. Khim., 14(7), 948 (1988).
31. Maria, R.M.C. and Santos, C.A. Thermochem. Acta, 125(15), 295 (1988).

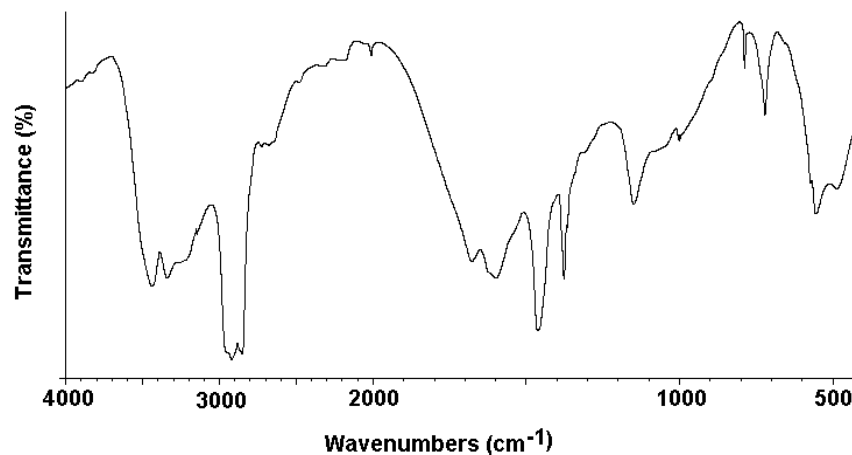
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**Table 2:** Infrared frequencies ( $\text{cm}^{-1}$ ) and assignments of the products obtained from the reaction of (a):  $\text{Co}(\text{HCOO})_2 \cdot x\text{H}_2\text{O}$  and (b):  $\text{Zn}(\text{HCOO})_2$  with urea at  $90^\circ\text{C}$  temperature.

Frequencies*, $\text{cm}^{-1}$		Assignments
$\text{CoCO}_3 \cdot x\text{H}_2\text{O}$	$\text{ZnCO}_3 \cdot x\text{H}_2\text{O}$	
3290, 3218	3433, 3331, 3208	$\nu(\text{O-H}); \text{H}_2\text{O}$
1571, 1356	1448, 1377	$\nu(\text{C-O}); \text{CO}_3^{2-}$
804, 752, 569	1162, 1080, 783	$\delta(\text{OCO}); \text{CO}_3^{2-}$

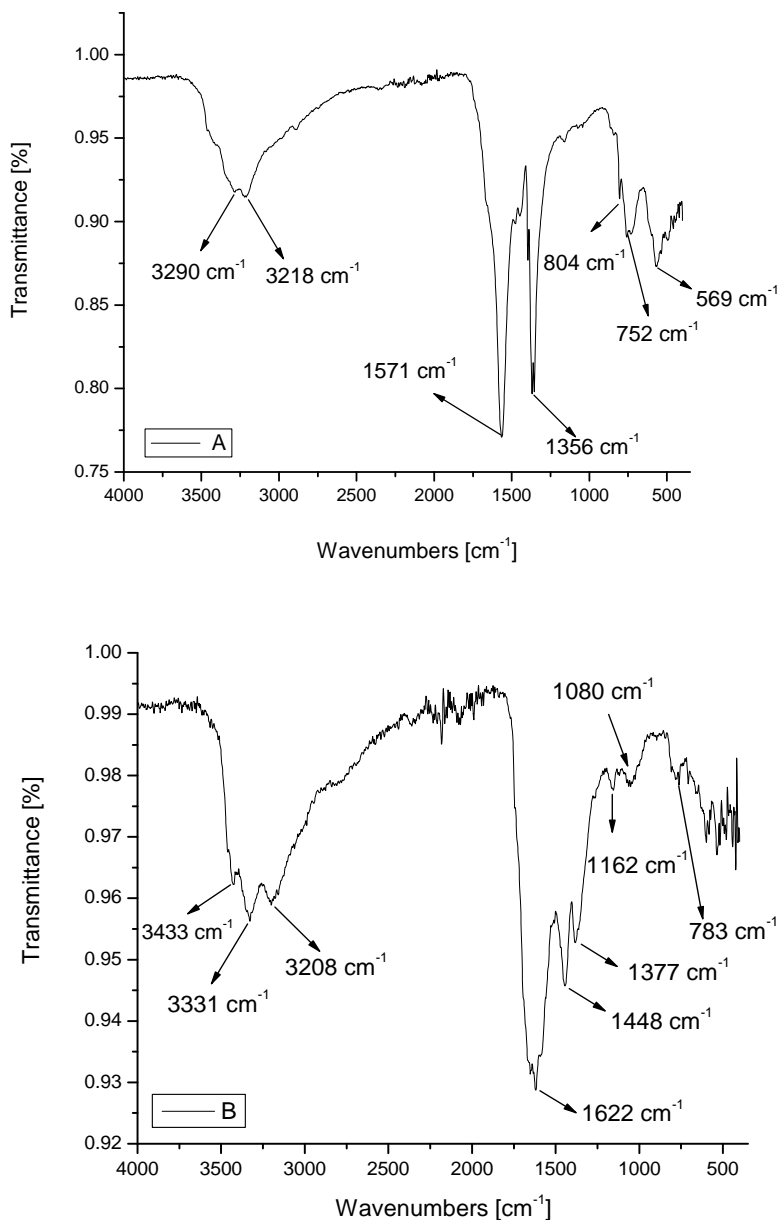


**Fig. 1:** Infrared spectrum of urea free ligand

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**Fig. 2:** Infrared spectra of  $\text{CoCO}_3 \cdot x\text{H}_2\text{O}$  (A) and  $\text{ZnCO}_3 \cdot x\text{H}_2\text{O}$  (B) compounds