

A Theoretical Study of Internal Pressure And Free Volume for Single Molecule of a Sample Liquid

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ABSTRACT: In the current research work a sample liquid is taken and attempt has been made for it to drive expressions for internal pressure and mean free volume occupied per molecules which depend up on the packing of molecules. Thermodynamical parameters like entropy, volume and temperature liquid property like viscosity are considered for deriving the expression of internal pressure. For the expression of free volume per molecule liquid properties like viscosity and the formula given by C.S.Suryanarayana & et.al are used. The free volume plays an important role in the propagation of ultrasonic waves in liquid.

I.INTRODUCTION

A sample liquid is taken and theoretical attempt has been made to derive expressions for internal pressure and free volume per molecule. The thermodynamic parameters sample like entropy, volume and temperature liquid properties like viscosity are considered for evaluation of expressions successfully.

II.THERMODYNAMICAL EQUATION STATE:

Thermodynamic importance of internal pressure is shown by bringing out its quantitative relationship with entropy and the partition function of the system. Internal pressure is shown to be related to the transport properties in liquids and solutions. The free energy of activation in the rate equations of Erring is almost the same as the cohesive energy of the system. In all rate processes, physical or chemical the energy barrier to be got over is just the cohesive energy barrier of the system in equilibrium.

The concept of free volume is a generalized aspect of the idea that its neighbors in a cell enclose each molecule. The free volume is broadly defined as the averages volume in which the center of the molecule can move inside the hypothetical cell due to the repulsion on the surrounding molecules.

The variation of entropy with volume and temperature may be represented in general terms as

At constant temperature, the first term is zero. Using Maxwell relation and for

$$ds = \left[\frac{\partial S}{\partial T} \right]_V dT + \left[\frac{\partial S}{\partial V} \right]_T dV \quad \text{-----(1)}$$

$$\left[\frac{\partial P}{\partial G} \right]_V = \frac{R}{V} \quad \text{-----(2)}$$

one mole of an ideal gas

So that $ds = R \ln V$ ----- (3)

And on integration we obtain for such an ideal gas,

$ST = R \ln V + \text{Constant}$ ----- (4)

If S is the value of S when the gas is at unit volume (at a given temperature) then

International Journal of Innovative Research in Science, Engineering and Technology

(An ISO 3297: 2007 Certified Organization)

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$$ST = S^{\circ} + R \ln V \quad \text{----- (5)}$$

Equ.(5) applies to a gas which is ideal. The problem now is to adapt it to liquids. This equ.(5) has a basic statistical interpretation in terms of the effect on the entropy of the larger volume available to each molecule as the pressure decreases. At constant temperature, the distribution of molecules among the possible vibration, rotational and electronic energy states is fixed and independent of pressure. So the only source of entropy change is the variation with the volume or pressure of the number of positions in space and the number of possible transnational energy levels opens to each molecules.

Statistically the entropy of an assembly of molecules forming thermodynamic systems is given by the equation.

$$S = K \ln W \quad \text{----- (6)}$$

Where W is the thermodynamic probability. W is the number of possible ways of distributing the molecules, among the various positions and energy levels available, which will lead to the observed state of the system. The entropy becomes less in a condensed phase as the free space is less than in a gas. Another important step in adapting equ.(5) to condensed phases consists in using for the volume V, the free volume Vf.

Newton and Eyeing [1] and also Len nard-Jones and Devon shire [2] expressed the entropy of vaporization as approximately

$$Sc = R \ln \left[\frac{Vg}{Vf} \right] \quad \text{----- (7)}$$

Where Vg is the volume of the vapour and Vf the free volume of the liquid. Vf is the characteristic property of a liquid.

A discussion of the importance of free volume in liquids by Rice [3] Franks [4] and Franks & Evans [5] and Hirschfelder et al. [6] shows the validity of introducing Vf in place of V in Equ.(5).

Hirsch elder et al. used the viral theorem to calculate the free volume of a molecule in a liquid. Assuming that the average force acting on a molecule is the sum of the internal pressure and the external pressure P, times the area of a face of the cube they obtained,

$$V_f = \left[\frac{kRT}{(\pi + p)(d + V_f^{1/3} / k)^2} \right]^3 \quad \text{----- (8)}$$

In Equ. (8), k is the proportionality factor, d the collision diameter of the molecules times a constant depending on the type of packing of the molecules. This relationship is very important since it shows that the free volume of a molecule at a particular temperature and pressure depends only on the internal pressure of liquid in which it is immersed. Hirsch elder et al. found that the proportionality factor k was 2 for normal liquids. Hence Eqn. (5) can be written as applicable to liquids. Hence Eqn. (5) can be written as applicable to liquids by replacing R by (Cp-Cv) as

$$ST = S^{\circ} - 3 (Cp-Cv) \ln \pi \quad \text{----- (9)}$$

The fundamental importance of this equation is very clear when the entropy is seen to be a function of internal pressure in a liquid. The constant S° will be the entropy at unit internal pressure πi is quantitatively related to the entropy of the medium by equ. (9) Even any microscopic changes occurring in the medium either due to molecular orientations or temperature changes should show up in a corresponding change in the internal pressure of the medium.

Internal pressure is more fundamental than the latent heat of vaporization and has the advantage that in multi-component systems it is the internal pressure of the entire system that counts, and is easily visualized in comparison with the latter which gets complicated when dealing with mixtures. It is obvious that the behavior of a is just related straight to the overall internal pressure of the medium. The molar free volume of a liquid in equilibrium with its vapor.

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Batschinsky[7] McLeod [8] Hildebrand [9] Hildebrand and Larnoreause [10] showed that viscosity, the most important intrinsic property of a liquid is quantitatively related to its free volume.

In the study of the thermodynamic properties of liquids, some of the properties are employed to derive the other. It is reasonable to represent the configuration integral in the form

$$QV = V_f N \times e^{-E/RT} \text{-----(10)}$$

Where V_f is called the free volumes, the significance of which is as follows. Actually E may be identified with the energy difference between the liquid and the gas in which there is no interaction at the absolute zero. V_f/N the free volume per molecule may be regarded as the effective volume accessible to the molecule in a liquid. In this volume, the potential is assumed to have a uniform value E , with a sharp rise to infinity at the boundary, the actual increase in potential is not so sudden as this would imply, but since repulsive forces decrease rapidly with distance, the approximations is not too serious.

The complete partition function if each molecule were confined within the cell, and were unable to escape will be,

$$Q_s = \left[\frac{(2\pi mkT)^{3/2}}{h^3} \cdot \frac{V_f}{N} \right] \left[q_i \exp\left(\frac{-E}{RT}\right) \right]^N \text{-----(11)}$$

Where q_i is the ideal partition function since each of the N molecules behave like an ideal gas molecule moving within its own volume V_f/N .

The complete partition function for a gas is

$$Q_e = \frac{1}{N!} \left[\frac{(2\pi mkT)^{3/2}}{h^3} \right]^N \cdot \int \dots \int \exp\left[\frac{-U}{kT}\right] dq_1, dq_2 \dots dq_{3N} \text{---(11.a)}$$

$\int \dots \int \exp[-U/kT] dq_1, dq_2 \dots dq_{3N}$ is called the configuration integral Q_n .

The configuration integral can be represented in an alternate form

$$Q_n = V_f N \exp(-E/RT) \text{-----(11.b)}$$

Where the significance of V_f will seen to be very important in liquids. However, if the molecules are not to regarded as localized elements, each completely confined to its own cell, it is necessary to introduce a factor $1/N!$.

Thus if the whole of the free volume V_f is available to all N molecules, the complete partition function for the system would be,

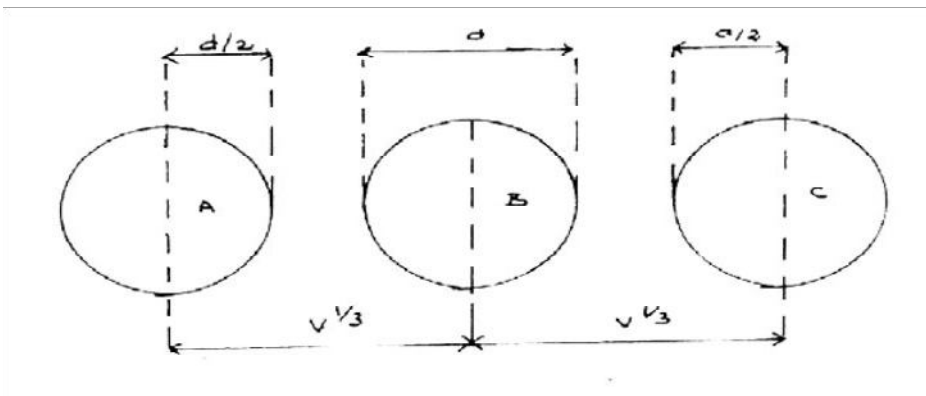
$$Q_l = \frac{1}{N!} \left[\frac{(2\pi mkT)^{3/2}}{h^3} \right]^N \cdot \left[\frac{V_f}{N} \right] \left[\left(q_i \exp\left\{ \frac{-E}{RT} \right\} \right) \right]^N \text{-----(12)}$$

Equ.(11) represents the partition function for solid and Eqn.(12) refers to that of a liquid. The increase in the entropy accompanying melting is merely due to the postulates that the whole of the free volume is available to each; molecule is the liquid state, whereas in the solid the molecule is restricted to its own cell. For this reason the entropy change per mole arrived at in this manner is called the communal entropy. The changes from solid to liquid does not involve a sudden transition from complete restriction to complete freedom of movement of the molecules accomplished by the introduction of the whole of the communal entropy.

III.DETERMINATION OF FREE VOLUME IN LIQUIDS:

For the calculation of free volume we adopt the method where the potential field is uniform within the cell but rising very sharply at the boundaries.

For simplicity, let us suppose that there is cubic packing of the molecules in the liquid, one molecule may be considered as oscillating about the origin and the six nearest neighbors are regarded as being fixed in their mean positions along these



the three axes. One of axes is represented in Fig.(1).

Figure.1.

If V is the mean volume per molecule in the liquid, then each molecule is at a distance $V^{1/3}$ from the origin. If d is the incompressible diameter of the molecule, $(2 V^{1/3} - d)$ is the distance, the central atom is free to move along each axis and the free volume per single molecule is given by

$$V_f = (2 V^{1/3} - d)^3 = 8 (V^{1/3} - d)^3 \text{ -----(13)}$$

For other types of molecular packing probably a similar equation will hold. Hence in general,

$$V_f = b^3 (V^{1/3} - d)^3 \text{ -----(14)}$$

Or

$$V_f = N b^3 (V^{1/3} - d)^3 \text{ -----(15)}$$

Where b is a constant whose value depends on packing. It takes a value of 2 for simple cubic structure and a value of 1.78 for face centered cubic structure and hexagonal closed pack structures.

The general equation relating the pressure of a liquid to the partition function is

Therefore, from Equ. (12)

$$P = kT \left[\frac{\partial \ln Q_1}{\partial V} \right]_T \text{ -----(16)}$$

(or)

$$P = NkT \left[\frac{\partial \ln V_f}{\partial V} \right]_T - \left[\frac{\partial E}{\partial V} \right]_T \text{ -----(17)}$$

Rearranging the term in Equ(17)& Substitute Equ. (14), we get,

$$V_f = \left[\frac{bRT}{P + \left(\frac{\partial E}{\partial T} \right)} \right]^3 \cdot \frac{1}{V^2} \text{ -----(18)}$$

International Journal of Innovative Research in Science, Engineering and Technology

(An ISO 3297: 2007 Certified Organization)

Vol. 2, Issue 12, December 2013

As far as liquids are concerned, the Equ. (18) has tremendous significance. Originally, U in Eqn. (11.a) meant a potential function depending on the configuration of the molecules. This U is replaced by a mean value E per mole, averaged over all accessible configurations. The internal degrees of freedom of molecules do not contribute to the potential function. Hence from the thermodynamic Eqn. (12) it is clear that the variation with respect to volume of the averaged potential function E, which is basically the potential energy of interactions at a given temperature, is replaceable in condensed phases by the variation of internal energy E with volume at the same temperature.

Where π_i is the internal pressure. As internal pressure values are of the order of 10³ atmospheres, P becomes negligible compared to $[\partial E / \partial V]_T$.

Hence, Eqn. (18) can be written as,

$$i.e. V_f \left[\frac{\partial E}{\partial V} \right]_{T, \pi_i} \approx \pi_i \frac{1}{V^2} \quad \text{-----(19)}$$

This equation has some unique characteristics. This is an important equation connecting quantitatively V_f and V in liquid systems. Again this gives a second confirmation of Hirschfelder et al [6] who first showed that the free volume of a molecule at a particular pressure of a liquid in which it is immersed.

In Eqn.(11) the molecules in solids are assumed to be localized whereas in Eqn. (12) giving the partition function for a liquid, the molecules are not localised but the entire free volume is available for them. Thus the free volume available in a liquid system may be called the communal free volume. Hence any change in the communal free volume V_f apart from situations analysed by Gurney [11,12] and Wood [13] will be an indication of change in the entropy of the system.

2. Internal Pressure, Free Volume and Sound Velocity:

The velocity of sound for liquids is 5 to 10 times greater than the average velocity predicted by the Kinetic theory of gases. Suppose there are three molecules A, B and C in a line, and the wave front travels from the inner edges of B with a velocity of sound U_g Applicable to an ideal gas, then

Where τ is the ratio of specific heats of the gas and M is its molecular weight. As A collides with B, however, the signal is transmitted almost instantaneously to the opposites edges of B.

$$U_g = \left[\frac{R_t \tau}{M} \right]^{1/2} \quad \text{-----(21)}$$

Thus although the wave front moves apparently through the distance $V/3$, it effectively travels a distance $V_f/3$ as seen in the above Fig.

In the case of gases $V_f/3$, is not too different from $V/3$ and in the case of solids $V_f/3$ is far too small compared to $V/3$. There are essential differences in the manner in which transverse stresses is propagated through gasses and solids. In a liquid, momentum is transferred partly by the diffusion as in gases, which is a relatively slow process, and partly by transverse elastic waves as in crystals, which is an extremely rapid process and the effective viscosity depends on their relative contributions. Thus the free volume plays an important role in ultrasonic propagation in liquids, and $V_f/3$ is very small in the case of solids. The relation between free volume and viscosity has been derived by Bingham [14] and McLeod [8].

Liquids are characterized by viscous relaxation when subjected to ultrasonic radiation. In other words, material response by way of round propagation to mechanical stress signified by ultrasonic irradiation is not instantaneous but the two are separated by a time interval, which is about the same as viscous relaxation time. Thermal relaxation is negligible during sound propagation in liquids.

C.V. Suryanarayana and Kuppusamy [15] derived a formula for free volume of liquid system by dimensional analysis. The ultrasonic velocity U in a liquids is a function of molecular weight M, viscosity η

$$U = k M^a \eta^b V_f^c \quad \text{-----(22)}$$

International Journal of Innovative Research in Science, Engineering and Technology

(An ISO 3297: 2007 Certified Organization)

Vol. 2, Issue 12, December 2013

Where k is a constant of proportionality.

Applying the method of dimensions. Hence Eqn. (22) becomes,

$$V_f = \left[\frac{Mu}{K\eta} \right]^{3/2} \text{-----(27)}$$

By substituting the known values of V_f for liquids, the value of k is found to be 4.28×10^9 . When M , U and V_f are given in c.g.s. Units and it is independent of the nature of liquids and temperature.

This equation is unique that the three fundamental bulk parameters of a liquid characterized its free volume.

From Eqn. (23) and Eqn.(27) we have

This equation is rigorous, excepting, for b , which depends on packing in the liquid system. Though there are uncertainties

$$\pi_i = bRT \left[\frac{K\eta}{U} \right]^{1/2} \cdot \frac{\rho^{2/3}}{M_{eff}^{7/6}} \text{-----(28)}$$

involved in determining the internal pressure of a liquid by other methods, Eqn.(28) is reliable. This is the first equation involving easily measurable quantities, which are fundamentally responsible for the structure of a liquid or a solution.

IV.CONCLUSION

Theoretical expressions for internal pressure and free volume per unit molecules have been successfully evaluated by considering thermodynamical properties like entropy, volume and temperature liquid property like viscosity. The values obtained for the internal pressure and free volume per molecule ion the present work coincide with values obtained in the literature available showing the validity of the present work.

ACKNOWLEDGEMENT

The authors are very thankful to Hon., Chairman, Chancellors, Vice-Chancellor, Director and Dean (Enge) and Dean R&D of Bharath University for their continuous support and encouragement given to bring out this paper.

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